9.9 Grignard Reagents

- Grignard reagents are prepared by the reaction of organic halides with magnesium turnings
 - An ether solvent is used because it forms a complex with the Grignard reagent which stabilizes it

$$\begin{array}{c} \textbf{RX} + \textbf{Mg} \xrightarrow{Et_2O} \textbf{RMgX} & \textbf{Grignard} \\ \textbf{ArX} + \textbf{Mg} \xrightarrow{Et_2O} \textbf{ArMgX} & \textbf{reagents} \end{array}$$

$$\frac{\text{CH}_{3}\text{I} + \text{Mg} \xrightarrow{\text{Et}_{2}\text{O}} \text{CH}_{3}\text{MgI}}{35^{\circ}\text{C}}$$

Methylmagnesium iodide (95%)

$$C_6H_5Br + Mg \xrightarrow{Et_2O} C_6H_5MgBr$$

Phenylmagnesium bromide (95%)

- Reaction of Grignard Reagents with Carbonyl Compounds
 - Nucleophilic attack of Grignard reagents at carbonyl carbons is the most important reaction of Grignard reagents
 - Reaction of Grignard reagents with aldehydes and ketones yields a new carbon-carbon bond and an alcohol

RMgX + C=0
$$\xrightarrow{(1) \text{ ether*}}$$
 R-C-O-H + MgX₂

Grignard Carbonyl reagent compound

Carbonyl Halomagnesium alkoxide

The strongly nucleophilic Grignard reagent uses its electron pair to form a bond to the carbon atom. One electron pair of the carbonyl group shifts out to the oxygen. This reaction is a nucleophilic addition to the carbonyl group, and it results in the formation of an alkoxide ion associated with Mg²⁺ and X⁻.

Step 2
$$\mathbb{R}$$
 $\stackrel{\circ}{\mathsf{C}}$ $\stackrel{\circ}{=}$ $\stackrel{\circ}{\mathsf{M}}$ $\stackrel{\circ}{\mathsf{M}}$ $\stackrel{+}{\mathsf{M}}$ \stackrel

Halomagnesium alkoxide

Alcohol

In the second step, the addition of aqueous HX causes protonation of the alkoxide ion; this leads to the formation of the alcohol and MgX₂.

Alcohols from Grignard Reagents

 Aldehydes and ketones react with Grignard reagents to yield different classes of alcohols depending on the starting carbonyl compound

- The Use of Lithium Reagents
 - Organolithium reagents react similarly to Grignard reagents
 - Organolithium reagents tend to be more reactive

The Use of Sodium Alkynides

 Sodium alkynides react with carbonyl compounds such as aldehydes and ketones to form new carboncarbon bonds

$$\mathbf{CH_3C} = \mathbf{CH} \xrightarrow{\mathrm{NaNH_2}} \mathbf{CH_3C} = \mathbf{CNa}$$

$$\mathbf{CH_3C} \stackrel{\delta-}{=} \stackrel{C}{\overset{\delta+}{\overset{C}{\longrightarrow}}} \stackrel{C}{\overset{C}{\longrightarrow}} \stackrel{C}{\longrightarrow} \stackrel{C}{\overset{C}{\longrightarrow}} \stackrel{C}{\overset{C}{\longrightarrow}} \stackrel{C}{\overset{C}{\longrightarrow}} \stackrel{C}{\overset{C}{\longrightarrow}} \stackrel{C}{\longrightarrow} \stackrel$$

Solved Problems

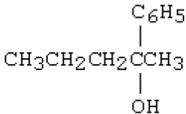
- 3. Which reagent(s) will distinguish between cyclopentanol and cyclopentane?
- A) Br₂/CCl₄
- B) KMnO₄ (cold)
- C) CrO₃/aqueous H₂SO₄
- D) NaOH (aq)
- E) A) and B)

21. What is the final product?

$$\begin{array}{ccc} D) & \text{CH}_3 \\ & | \\ & \text{CH}_3\text{CC}{\equiv}\text{CONa} \\ & | \\ & \text{CH}_3 \end{array}$$

E) None of these

27.



Your task is to synthesize OH through a Grignard synthesis. Which pairs of compounds listed below would you choose as starting materials?

A)

$$\parallel$$
 CH₃CH₂CH₂Br and CH₃CC₆H₅

B)

 \parallel CH₃CH₂CH₂CH and C₆H₅Br

C)

- D) More than one of these
- E) None of these

38. Which of the following would serve as a synthesis of racemic:

II
$$\longleftrightarrow$$
 CH₂CCH₃ + CH₃CH₂MgBr \longleftrightarrow (1) Et₂O \longleftrightarrow NH₄+

|| CH₂CCH₂CH₃ + CH₃MgI
$$\frac{(1) \text{ Et}_2\text{O}}{(2) \text{ NH}_4^+}$$

- A) I
- B) II
- C) III
- D) All of the above
- E) None of the above